

Evaluation of effectiveness of a 'Constructive approach of teaching' in learning and application of chemical formulae of ionic compounds among the 9th grade students

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Author: Darshana Kamat

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S. B. Patil Public School

Abstract

The objective of this research is to test the effect of a constructive approach towards students' understanding in writing the chemical formulae of ionic compounds. This quasi-experimental case study design research involves 79 students divided into two groups: one an experimental group (n = 41) and the other a control group (n = 38). The findings revealed that there were statistically significant differences between the experimental and the control groups, in favor of the experimental group. Our findings are based on both quantitative data (performance of students in tests) and qualitative data (collected through students' work). The findings proved that the constructive approach of teaching is more effective in assisting students to write and apply the chemical formula of ionic compounds. Students also have a positive perception towards the usage of constructive approach in teaching and learning activity. This study implies that the elements of experiential constructivism should be infused during the planning of teaching and learning to ensure that learning is more heuristic.

Introduction:

Mastering the symbolic representation is essential to understand the chemistry concepts and it is fundamental in writing the chemical equations. Students fail to explore chemistry when they face difficulties in writing chemical equations which arise from their inability to write formulae of chemical compounds correctly.

Although students are able to manipulate chemical formulae, they tend to solve it as a mathematical problem without understanding the concept behind it. So facing problems in retention of knowledge and use of it in writing the chemical equations.

Research Background:

The understanding of valency, appreciation of the concept of polyatomic ions and molecules

and ultimately the production of correct chemical formula depends on students' knowledge of bonding. Unfortunately, concepts in chemical bonding are highly abstract. Hence, a suitable pedagogical design is needed in visualizing the concepts involved in writing formulas of compounds. The researcher implemented the proposed action strategies by interventions to improve student's performance through designing student-centered constructivist approaches that help easily to avoid errors in writing the basic elements of Chemistry Language thereby developing confidence in how to easily learn chemistry concepts such as writing chemical formulae.

Objective:

- 1.To analyze the positive impact of new approaches in teaching and learning on students' understanding and achievement while increasing their involvement.
2. To analyze the positive impact of new approaches in teaching and learning on students' application of chemical formulae of ionic compounds

Research questions:

Is the constructive approach used as a learning tool in writing the chemical formula of ionic compounds more effective in increasing the students' knowledge, understanding and application than conventional teaching methods?

Research Hypothesis:

After the overall interventions and use of constructive approach of teaching, there will be significant improvement in knowledge, understanding and application of chemical formulae of ionic compounds among the 9th grade students.

Null () Hypothesis:

After the overall interventions and use of constructive approach of teaching, there will not be any significant improvement in knowledge, understanding and application of chemical formulae of ionic compounds among the 9th grade students.

Time Duration: 2 Months

- Preparation of action plan - 2 weeks
- Drafting of Pre-test, Post Test and Questionnaire Checklist - 2 weeks

- Interventions – 3 periods of 40 min each
- Administration of Post-test and the questionnaire-checklist to the respondent 2 periods.
- Data analysis, calculation and tabulation of result- 2 weeks
- Completion of research - 2 week

Population and Sample:

Population - Students of grade 9 from S. B. Patil Public School, Ravet affiliated to CBSE (Central Board of Secondary Education)

Sample -

41 students of grade 9 division E (experimental group) & 38 students of grade 9 division F (controlled group)

Research Design: Quasi- experimental case study design

Teacher compared the two group's (controlled and experimental group) performance on the pre and post questionnaire which will determine if one group performs significantly better than the other on the assessment based on pedagogy (constructivist classroom vs. traditional). Knowing what works best for students in the classroom will allow the researcher to improve teaching skills and will likely increase learning for students.

Variables of Study:

- **An Independent variable** - classroom pedagogy (constructivist classroom vs. traditional).
- **Dependent variables** - Pre-Test and Post-Test scores.

Tools of Data collection:

1. Pre-test and Post Test scores
2. Students Worksheets

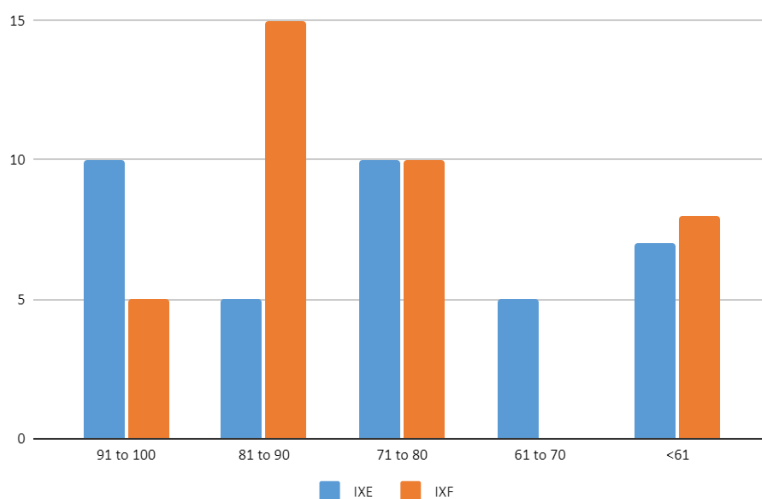
Two sections of grade 9 were enrolled in the study. The classes were selected randomly.

Table1 Number of students in each study group

Study Group	Grade & Div	No of students
Control group	9 F	38
Experimental group	9 E	41

Pretest- A pretest of 20 marks on the topics atoms, ions, molecules, valency was administered to both the experimental (IX E) and control group (IX F) as before their study of the chemical formulae to ascertain their prior knowledge on the topic. And to check if there is any significant difference between the two groups with regard to their academic achievement in the above topics.

Quantitative Analysis of Pre test		
	No. of students	No. of students
Criteria (%)	IXE	IXF
91 to 100	10	5
81 to 90	5	15
71 to 80	10	10
61 to 70	5	0
<61	7	8
	2 Students Absent	1 student Absent



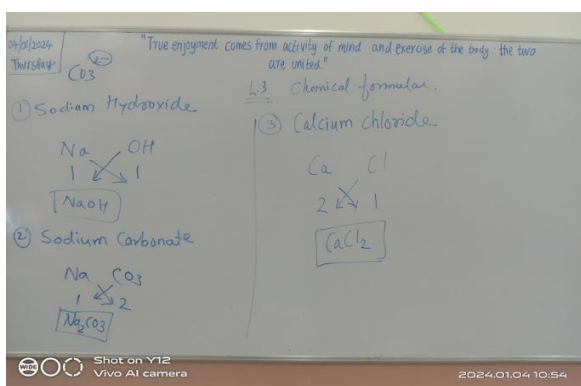
Tools of Data Analysis: Mean of Pre and Post test scores

Tools of Data Representation: Tables, Graphs, Charts

Interventions:

I) Traditional Method of Teaching- (For control group)

The students in the Control Group were instructed only with traditionally designed learning material. Textbook was used as a primary material. Most of the time, the teacher presented the topics and the students listened to their teacher and answered the questions asked by their teacher. Teacher explained many examples of chemical formulae. At the same time, they carried out activities in their text-books, solved some examples as homework practice.



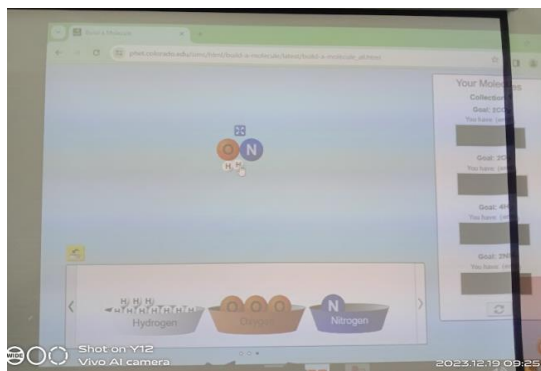
II) Constructivist Method of teaching -

The constructivist classroom model suggests that constructivist lessons should **engage** students, allow them to **explore**, aid them in **explaining** their experience, learning is **elaborated**, and the lesson includes **evaluation**.

Five “E” Model -

1. Engage - https://phet.colorado.edu/sims/html/build-a-molecule/latest/build-a-molecule_all.html

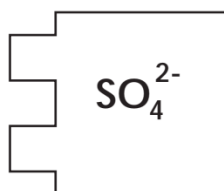
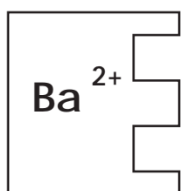
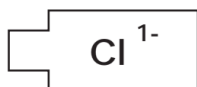
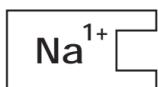
Teacher used the Phet simulation to reinforce the students' understanding of ionic molecules. This simulation allows students to visualize how different molecules and compounds form. This activity is found to be an ideal way to engage students and ascertain their prior knowledge of atoms and molecules.



2. Explore - A hands-on activity called Jigsaw puzzles was introduced at this point.

Puzzles that represent the positive and negative ions were built and the formulae of the ions were written on the puzzles.

In the jigsaw puzzle game, students tried to solve the puzzle to form a rectangular shape figure. When students combined the puzzles with loops and heads to solve the puzzles, they visualized the combination of positive and negative ions to form neutral compounds.



i) Activity one- Make the given compounds using the appropriate pieces:

Example-

Sodium Chloride formula _____

Potassium Bromide formula _____

Ammonium Nitrate formula _____

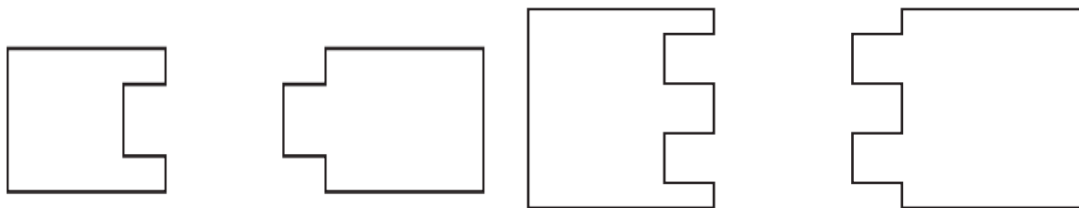


ii) Activity Two - Make some compounds of your own choosing and write the formulas:

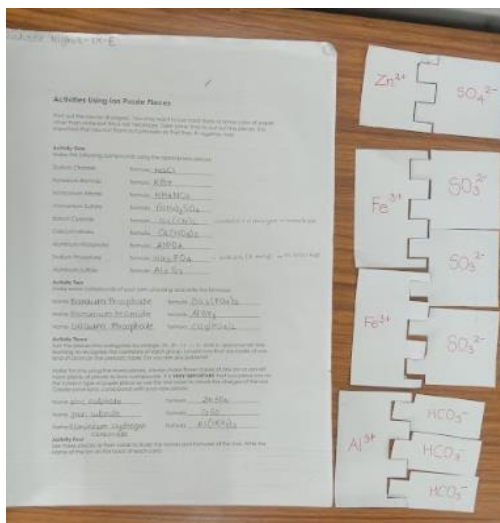
Name _____ formula _____
 Name _____ formula _____

iii) Activity Three - Make the compounds using blank pieces.

Ask students to use blank pieces and make ions. Always make three copies of any ion so students will have plenty of pieces to form compounds. It is very important that students place ions on the correct type of puzzle piece so use the textbook to check the charges of the ions. Create some ionic compounds with the new pieces:



Name _____ formula _____
 Name _____ formula _____



3. **Explain** - Students were encouraged to explain their own created ionic compounds with the new pieces in their own words, asked for evidence and clarification of their explanation, and listen critically to one another's explanation. Students used observations and recordings in their explanations. At this stage researcher provided definitions and explanations using students' previous experiences as a basis for this discussion

4. Elaborate -

i) A LEGO - based interlocking building block Activity.

During this activity student were encouraged to apply concepts and skills in new (but similar) situations of building a compound with LEGO blocks.



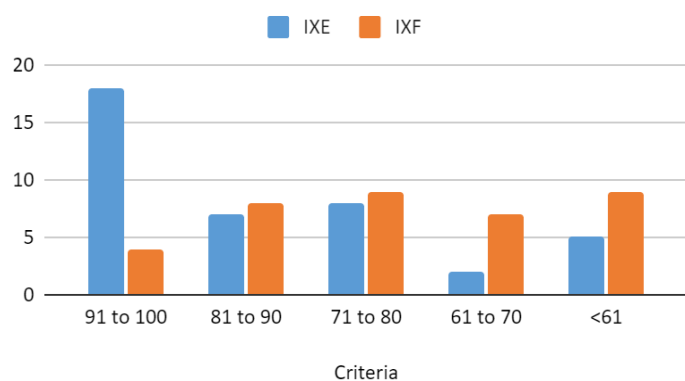
ii) Students were energized to apply the learned Chemical formulae to be used in chemical equations to describe chemical reactions, reactants and products.

5. **Evaluate** - Students cross checked their worksheets with other group students

Data Collection of Post test after Interventions-

Quantitative Analysis of Post test		
	No. of students	No. of students
Criteria	IXE	IXF
91 to 100	18	4
81 to 90	7	8
71 to 80	8	9
61 to 70	2	7
<61	5	9
	1 Student Absent	2 Students Absent

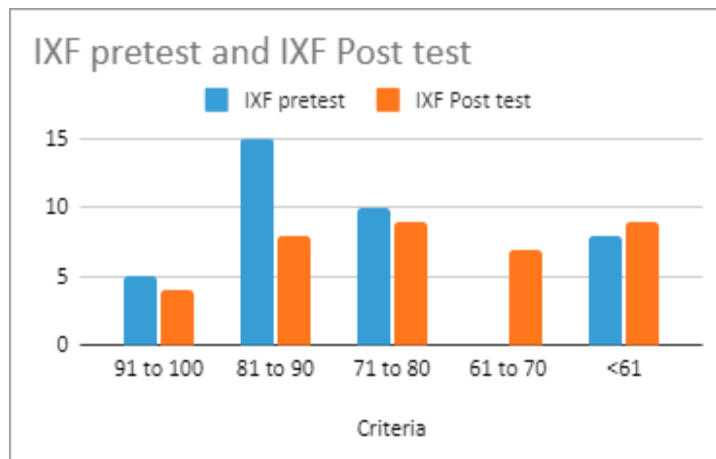
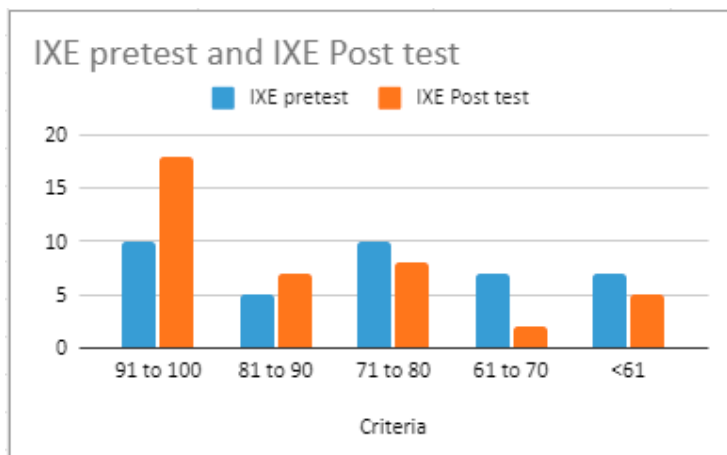
IXE and IXF



The following quantitative data compares the students' scores of pre and posttests.

Criteria	IXE pretest	IXE Post test
91 to 100	10	18
81 to 90	5	7
71 to 80	10	8
61 to 70	7	2
<61	7	5

Criteria	IXF pretest	IXF Post test
91 to 100	5	4
81 to 90	15	8
71 to 80	10	9
61 to 70	0	7
<61	8	9



Data Analysis:

1. Students in the Experimental group (IX E), who taught using a constructive method, showed a statistically significant increase in score on the posttest compared to the pretest, with a significant difference in outcomes between the two tests.
2. Students in the control group (IX F), who were taught using conventional methods, were unable to meaningfully answer the application-based problems. Consequently, demonstrating less development in the posttest.
3. Thus, the results reject the null hypothesis and accept the Research hypothesis to prove that after the overall interventions and use of constructive approach of teaching, there will be significant improvement in knowledge, understanding and application of chemical formulae of ionic compounds among the 9th grade students.

Major Findings:

1. This suggests that the Constructive method of instruction is useful in helping students grasp how to compose ionic complex chemical formulas. Visualization is therefore crucial to comprehending the ideas underlying the chemical representation.
2. The results demonstrate the potential potency and efficacy of a novel method for composing an ionic compound's chemical formula.
3. As a result, the experiential learning activity known as the constructive method of teaching may offer an engaging classroom setting that will pique students' curiosity and enthusiasm.

The findings of this study provide additional evidence for the value of experiential learning in piquing students' interest in picking up new concepts.

Conclusion - The constructivism approach gives students the opportunity to develop new knowledge from their visualization in writing chemical formulas of ionic compounds. From the students' feedback we can conclude that the innovative approach in writing the chemical formula of ionic compounds is accepted by students. The findings of this study support the positive impact of constructive approaches in teaching and learning on students' understanding and achievement while increasing their involvement.

In conclusion, the teaching and learning using the constructive approach in the topic of chemical formula assists the early learners in visualizing, understanding and applying the main concept underlies in writing the chemical formula of ionic compounds. Thus, preventing them from memorizing and accepting blindly formulas.

Take away-

This research implies that this constructive approach should be implemented in teaching and learning to improve the students' understanding in chemistry. For this, teachers should be well comprehended on the design and usage of the constructive approach as they play the main role of facilitators in teaching and learning. Finally, the insights gained from this research are expected to inculcate heuristic as well as thoughtful learning among the young chemistry learners.

Significance of the study: -

This research would be of great help for the teachers who would use the innovative techniques in their teaching for giving a good learning experience to the students.

Recommendation: -

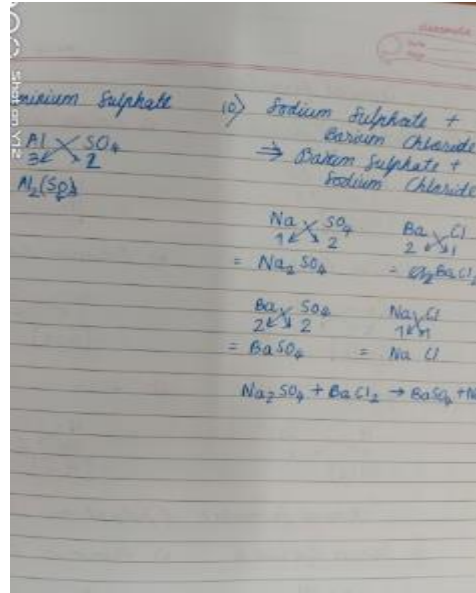
Henceforth, the same strategy could be used for the 9th Graders to enhance their visualization, understanding and application of chemical formulae.

References

1. Chaille, C. (2008). Constructivism across the curriculum: Big ideas as inspiration. Upper Saddle River, NJ: Allyn & Bacon
2. Constructivism in the Classroom-
<https://edpsych.pressbooks.sunycreate.cloud/chapter/constructivism-in-the-classroom/>
3. Constructivism: learning theories and approaches to research- Anna MacLeod

Annexure 1

Traditional Method of Teaching-



Annexure 2

Worksheet of jigsaw puzzles

Handwritten student work on a worksheet:

Activities Using Ion Puzzle Pieces

Activity One

Write the following compounds using the appropriate prefixes:

Sodium Chloride	Formula	<u>NaCl</u>
Potassium Bromide	Formula	<u>KBr</u>
Ammonium Nitrate	Formula	<u>NH₄NO₃</u>
Ammonium Sulphate	Formula	<u>(NH₄)₂SO₄</u>
Silver Chloride	Formula	<u>AgCl</u>
Calcium Nitrate	Formula	<u>Ca(NO₃)₂</u>
Aluminium Phosphate	Formula	<u>AlPO₄</u>
Sodium Phosphate	Formula	<u>Na₃PO₄</u>
Aluminium Sulfate	Formula	<u>Al₂(SO₄)₃</u>

Activity Two

Write some compounds of your own choosing and write the formula:

Name	<u>Calcium phosphate</u>	Formula	<u>Ca₃(PO₄)₂</u>
Name	<u>Aluminium nitrate</u>	Formula	<u>Al(NO₃)₃</u>
Name	<u>Sodium carbonate</u>	Formula	<u>Na₂CO₃</u>

Activity Three

Sort the pieces into categories for cations: 1a, 1b, 1c, 2a, 2b, and 3. Record some that starting to recognize the members of each group. Locate ions that are made of one kind of atom on the periodic table. Do you see any patterns?

Write for ions using the blank pieces. Always make three copies of any ion so you will have plenty of pieces to form compounds. **VERY IMPORTANT!** Not only place ions at the correct type of ionic charge on the left, but check to check the charge of the ions. Double check ionic compounds with your new pieces.

Name	<u>Aluminium oxide</u>	Formula	<u>Al₂O₃</u>
Name	<u>Hydroxide</u>	Formula	<u>OH⁻</u>
Name	<u>Aluminium phosphate</u>	Formula	<u>AlPO₄</u>

Activity Four

Use these pieces as both cations to build the ionic and formula of the ions. Write the name of the ion on the back of each card.

Handwritten student work on a worksheet with jigsaw puzzle pieces:

Activities Using Ion Puzzle Pieces

Activity One

Write the following compounds using the appropriate prefixes:

Sodium Chloride	Formula	<u>NaCl</u>
Potassium Bromide	Formula	<u>KBr</u>
Ammonium Nitrate	Formula	<u>NH₄NO₃</u>
Ammonium Sulphate	Formula	<u>(NH₄)₂SO₄</u>
Silver Chloride	Formula	<u>AgCl</u>
Calcium Nitrate	Formula	<u>Ca(NO₃)₂</u>
Aluminium Phosphate	Formula	<u>AlPO₄</u>
Sodium Phosphate	Formula	<u>Na₃PO₄</u>
Aluminium Sulfate	Formula	<u>Al₂(SO₄)₃</u>

Activity Two

Write some compounds of your own choosing and write the formula:

Name	<u>Calcium phosphate</u>	Formula	<u>Ca₃(PO₄)₂</u>
Name	<u>Aluminium nitrate</u>	Formula	<u>Al(NO₃)₃</u>
Name	<u>Sodium carbonate</u>	Formula	<u>Na₂CO₃</u>

Activity Three

Sort the pieces into categories for cations: 1a, 1b, 1c, 2a, 2b, and 3. Record some that starting to recognize the members of each group. Locate ions that are made of one kind of atom on the periodic table. Do you see any patterns?

Write for ions using the blank pieces. Always make three copies of any ion so you will have plenty of pieces to form compounds. **VERY IMPORTANT!** Not only place ions at the correct type of ionic charge on the left, but check to check the charge of the ions. Double check ionic compounds with your new pieces.

Name	<u>Aluminium oxide</u>	Formula	<u>Al₂O₃</u>
Name	<u>Hydroxide</u>	Formula	<u>OH⁻</u>
Name	<u>Aluminium phosphate</u>	Formula	<u>AlPO₄</u>

Activity Four

Use these pieces as both cations to build the ionic and formula of the ions. Write the name of the ion on the back of each card.

Jigsaw puzzle pieces showing ions:

- Zn²⁺ and SO₄²⁻
- Fe³⁺ and SO₄²⁻
- Fe³⁺ and SO₃²⁻
- Fe³⁺ and SO₃²⁻
- Al³⁺ and HCO₃⁻
- Al³⁺ and HCO₃⁻
- Al³⁺ and HCO₃⁻

Annexure 3

Answer sheet of Pre-test

Pre-Test Questionnaire
 S.B. PATIL PUBLIC SCHOOL
 S.No. 131, 1st Flr, Road, Pune-412201
 Action Research (A.Y. 2023-24)
 Pre-Test Questionnaire

Sub: Chemistry
 Std: IX
 Name of the Student: Bhaisavi Mahesh Gawde
 Date: 16/12/23
 Max. Marks: 20
 Time: 30 min

- Write the symbols of
 a) Iron Fe b) Sulphur S c) Calcium Ca d) Aluminium Al (2M)
- Write the electronic configuration of an element whose atomic number is 12. 2, 8, 2 (1M)
- Predict the valency of the following elements
 (a) A (Atomic number 5) = 3
 (b) C (Atomic number 14) = 4 (1M)
- Identify the Anion from the following figures. What is the valency of sodium atom? Give reason. (2M)

The valency of sodium atom is 1. The sodium atom has 1 electron in its outer shell, which is the combining capacity of sodium.

5. Complete the following gaps in the given table: (3M)

Element	Protons	Electrons	Atomic Number	Mass Number
A	8	8	8	16
B	11	10	11	23
C	1	1	1	2
D	11	11	11	23

- Give one point of difference between an atom and an ion.
An atom is electrically neutral but ion is a charged particle.
- Phosphorus is electrically neutral but ion is a charged particle.
- The electronic configuration of phosphorus atom is 2, 8, 5. Give the electronic configuration of P³⁻ ion. 2, 8, 10, 3 (1M)
- The atomic number of Al is 13. Which will be the number of electrons in Al³⁺? 10 (1M)
- Give one example each of a polyatomic cation and an anion.
Ammonium Carbonate = cation, Hydroxide OH⁻ = Anion (2M)
- Write the cations and anions present in the following compounds:
 (a) CH_{3COOH} = CH_{3COO}⁻ and H⁺ (1M)
 (b) NaCl = Na⁺ and Cl⁻ (1M)
- State the number of atoms present in each of the following chemical species
 (a) CO₂ = 4 (b) P₂O₅ = 7 (2M)
- Complete the table on the basis of information available in the symbols given below. (3M)

Element	Protons	Electrons	Atomic Number	Mass Number
Calcium	20	20	20	40
Carbon	6	6	6	12
Bromine	35	35	35	80

Pre-Test Questionnaire
 S.B. PATIL PUBLIC SCHOOL
 S.No. 131, 1st Flr, Road, Pune-412201
 Action Research (A.Y. 2023-24)
 Pre-Test Questionnaire

Sub: Chemistry
 Std: IX
 Name of the Student: Bhaisavi Mahesh Gawde
 Date: 16/12/23
 Max. Marks: 20
 Time: 30 min

- Write the symbols of
 a) Iron Fe b) Sulphur S c) Calcium Ca d) Aluminium Al (2M)
- Write the electronic configuration of an element whose atomic number is 12. 2, 8, 2 (1M)
- Predict the valency of the following elements
 (a) A (Atomic number 5) = 3
 (b) C (Atomic number 14) = 4 (1M)
- Identify the Anion from the following figures. What is the valency of sodium atom? Give reason. (2M)

The valency of sodium atom is 1. The sodium atom has 1 electron in its outer shell, which is the combining capacity of sodium.

5. Complete the following gaps in the given table: (3M)

Element	Protons	Electrons	Atomic Number	Mass Number
A	8	8	8	16
B	11	10	11	23
C	1	1	1	2
D	11	11	11	23

- Give one point of difference between an atom and an ion.
An atom is electrically neutral but ion is a charged particle.
- Phosphorus is electrically neutral but ion is a charged particle.
- The electronic configuration of phosphorus atom is 2, 8, 5. Give the electronic configuration of P³⁻ ion. 2, 8, 10, 3 (1M)
- The atomic number of Al is 13. Which will be the number of electrons in Al³⁺? 10 (1M)
- Give one example each of a polyatomic cation and an anion.
Ammonium Carbonate = cation, Hydroxide OH⁻ = Anion (2M)
- Write the cations and anions present in the following compounds:
 (a) CH_{3COOH} = CH_{3COO}⁻ and H⁺ (1M)
 (b) NaCl = Na⁺ and Cl⁻ (1M)
- State the number of atoms present in each of the following chemical species
 (a) CO₂ = 4 (b) P₂O₅ = 7 (2M)
- Complete the table on the basis of information available in the symbols given below. (3M)

Element	Protons	Electrons	Atomic Number	Mass Number
Calcium	20	20	20	40
Carbon	6	6	6	12
Bromine	35	35	35	80

Answer sheet of Post-test

Post-Test Questionnaire
 S.B. PATIL PUBLIC SCHOOL
 S.No. 131, 1st Flr, Road, Pune-412201
 Action Research (A.Y. 2023-24)
 Post-Test Questionnaire

Sub: Chemistry
 Std: IX
 Name of the Student: Vidushi Venusa
 Date: 16.01.2024
 Max. Marks: 20
 Time: 30 min

- Write the formula and names of the compounds formed by the following ions. (3M)
- a) Potassium ion and iodide ion - KI Potassium iodide
 b) Sodium ion and sulphide ion - Na₂S Sodium sulphide/sulphide
 c) Aluminium ion and phosphate ion - AlPO₄ Aluminium phosphate/phosphate
- Name the cation and anion which constitute the molecule of Calcium carbonate. Ca²⁺ and CO₃²⁻ (1M)
- Write down the names of compounds represented by the following formulae: (4M)
- a) Al₂(SO₄)₃ - Aluminium sulphate b) CaCl₂ - Calcium chloride
 c) K₂SO₄ - Potassium sulphate d) KNO₃ - Potassium nitrate
- Give the chemical formula for the following compounds: (3M)
- a) Hydrogen chloride - HCl b) Aluminium fluoride - AlF₃ c) Magnesium sulphide - MgS
 d) Sodium oxide - Na₂O e) Aluminium chloride - AlCl₃ f) Sodium sulphide - Na₂S
- Magnesium hydroxide - Mg(OH)₂ i) Strontium carbonate - Li₂CO₃ j) barium hydroxide - Ba(OH)₂
- magnesium phosphate - Mg₃(PO₄)₂
- In a common experiment in the general chemistry laboratory, magnesium metal is heated in the presence of oxygen to produce Magnesium oxide powder. This Magnesium oxide when dissolved in water produces Magnesium hydroxide. Translate the given information in chemical equation. (2M)
- Write the word equations below as chemical equations. (5M)
- (a) Sodium phosphate and calcium chloride react to form calcium phosphate and sodium chloride.
 $Na_3PO_4 + CaCl_2 \rightarrow Ca_3(PO_4)_2 + NaCl$
 (b) Aluminium and hydrochloric acid react to form aluminium chloride and hydrogen gas.
 $Al + HCl \rightarrow AlCl_3 + H_2$
 (c) Zinc and lead (II) nitrate react to form zinc nitrate and lead.
 $Zn + Pb(NO_3)_2 \rightarrow Zn(NO_3)_2 + Pb$
 (d) Magnesium chloride and sodium hydroxide react to produce magnesium hydroxide and sodium chloride.
 $MgCl_2 + NaOH \rightarrow Mg(OH)_2 + NaCl$
 (e) Solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas.
 $CaCO_3 \xrightarrow{heat} CaO + CO_2(g)$

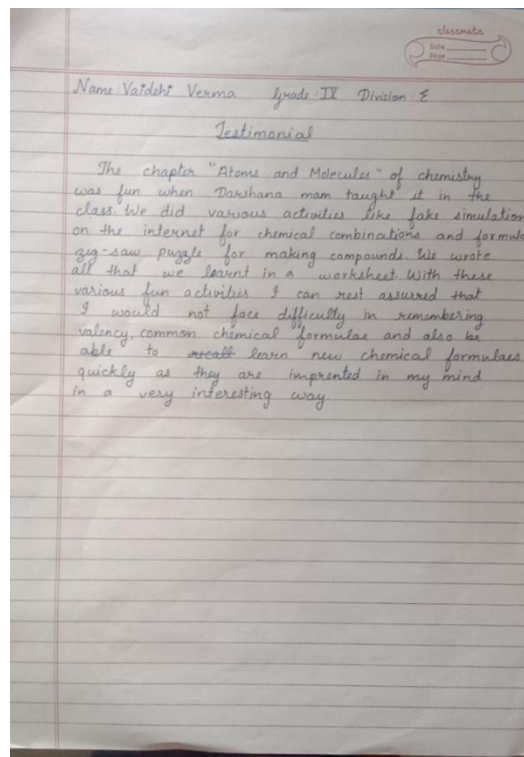
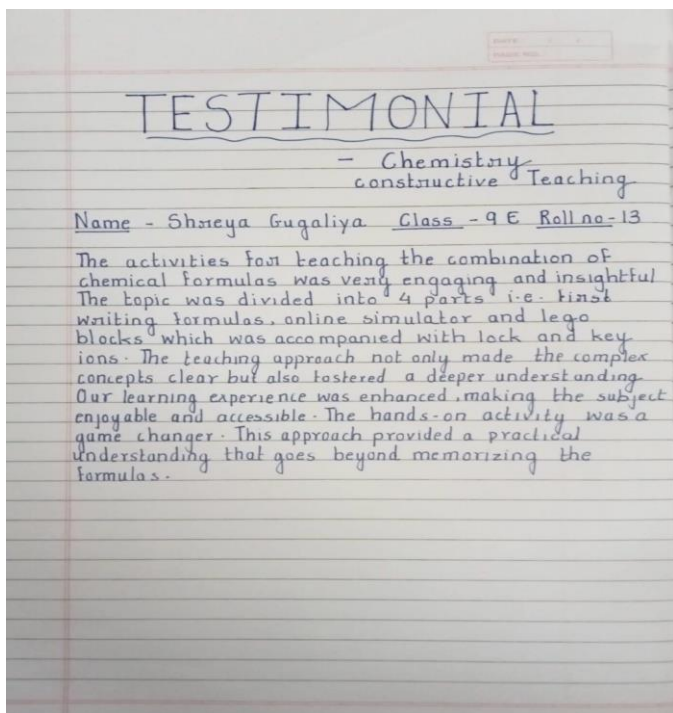
Post-Test Questionnaire
 S.B. PATIL PUBLIC SCHOOL
 S.No. 131, 1st Flr, Road, Pune-412201
 Action Research (A.Y. 2023-24)
 Post-Test Questionnaire



Sub: Chemistry
 Std: IX
 Name of the Student: Rohit Sagar Chavhan
 Date: 16/12/2024
 Max. Marks: 20
 Time: 30 min

- Write the formula and names of the compounds formed by the following ions. (3M)
- a) Potassium ion and iodide ion - KI Potassium iodide
 b) Sodium ion and sulphide ion - Na₂S Sodium sulphide and iodide
 c) Aluminium ion and phosphate ion - AlPO₄ Aluminium sulphate
- Name the cation and anion which constitute the molecule of Calcium carbonate. Ca²⁺ and CO₃²⁻ (1M)
- Write down the names of compounds represented by the following formulae: (4M)
- a) Al₂(SO₄)₃ - Aluminium sulphate b) CaCl₂ - Calcium chloride
 c) K₂SO₄ - Potassium sulphate d) KNO₃ - Potassium nitrate
- Give the chemical formula for the following compounds: (5M)
- a) Hydrogen chloride - HCl b) Aluminium fluoride - AlF₃ c) Magnesium sulphide - MgS
 d) Sodium oxide - Na₂O e) Aluminium chloride - AlCl₃ f) Sodium sulphide - Na₂S
- Magnesium hydroxide - Mg(OH)₂ i) Strontium carbonate - Li₂CO₃ j) barium hydroxide - Ba(OH)₂
- magnesium phosphate - Mg₃(PO₄)₂
- In a common experiment in the general chemistry laboratory, magnesium metal is heated in the presence of oxygen to produce Magnesium oxide powder. This Magnesium oxide when dissolved in water produces Magnesium hydroxide. Translate the given information in chemical equation. (2M)
- Write the word equations below as chemical equations. (5M)
- (a) Sodium phosphate and calcium chloride react to form calcium phosphate and sodium chloride.
 $Na_3PO_4 + CaCl_2 \rightarrow Ca_3(PO_4)_2 + NaCl$
 (b) Aluminium and hydrochloric acid react to form aluminium chloride and hydrogen gas.
 $Al + HCl \rightarrow AlCl_3 + H_2$
 (c) Zinc and lead (II) nitrate react to form zinc nitrate and lead.
 $Zn + Pb(NO_3)_2 \rightarrow Zn(NO_3)_2 + Pb$
 (d) Magnesium chloride and sodium hydroxide react to produce magnesium hydroxide and sodium chloride.
 $MgCl_2 + NaOH \rightarrow Mg(OH)_2 + NaCl$
 (e) Solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas.
 $CaCO_3 \xrightarrow{heat} CaO + CO_2(g)$

Annexure 4

Students Reflection



Indicators for teaching effectiveness  



Questions Responses **18** Settings

Total points: 0

no

It was a great activity, it helped me in understanding about the chemical formula topic as it was very interesting and engaging activity. More such activities should be conducted by the school.

The topic first introduced by the activity form and then encountered with solving problems helped better understanding rather than passively memorizing formulae.

No

Explained very perfectly in fact I don't have to take stress about chemistry

The activity was very interesting and got to know how to make these chemical formulae by using the blocks of the elements.

It was interesting activity. I understood that hard concept very easily, now I can solve questions related to topic Chemical formulae

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